Topic Reminder Q5
Ionic bonding
#3-1, 2

Provide the name and or formula for the following questions

1. $Mg(OH)_2$ = Magnesium Hydroxide
2. $FeCl_2$ = Iron (II) Chloride
3. $Ca(OH)_2$= Calcium Hydroxide
4. Magnesium Hydroxide = Magnesium Hydroxide
5. Iron(III) sulfate = $Fe_2(SO_4)_3$
6. Calcium acetate = $Ca(C_2H_3O_2)_2$

Properties of ionic compounds listed above,

7. Which substances has the highest melting point?

$Fe_2(SO_4)_3$, highest charges $+2$ -2

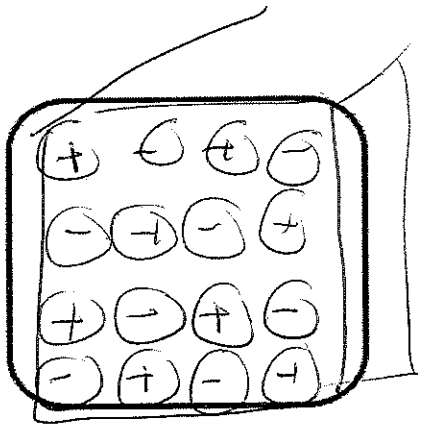
8. Create a substance that would have a lower melting point then all the substances above.

NaCl

9. Draw a picture of the substance in #8.

$Na^+ = +$

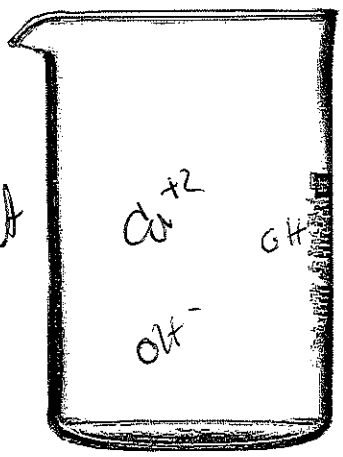
$Cl^- = -$



10. Write out the reaction demonstrating the dissolving of $Ca(OH)_2$.

$Ca(OH)_2 \rightarrow Ca^{+2} + 2OH^-$

11. Draw a picture of $Ca(OH)_2$ in the beaker below.



you can have more particles but always 1:2 ratio of Ca OH